



Explain, in terms of attractive forces between particles, why  $\text{LiCl}$  is a solid at room temperature but  $\text{NCl}_3$  is a liquid with a relatively low boiling point.

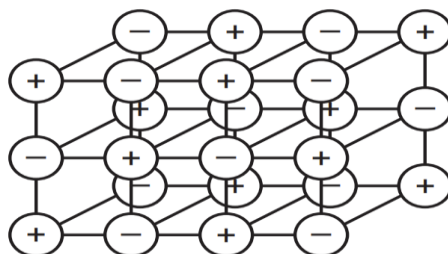
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4. The structure of a typical ionic compound is a regular arrangement of positive negative ions.



- (i) What is the name of this regular arrangement of particles?

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- (ii) Give **two** physical properties of ionic compounds.

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- (iii) Explain the term ionic lattice.

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